

Characteristics let's
move to the principles
of the



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Characteristics of equilibrium reaction are different in some ways. Now, we're going to talk about one thing that is not going to change. So, what characteristics define a state? In a chemical reaction, is the state in which both reactants and products are presented in concentrations which have no further tendency to be changed with time so that there is no observable change in the properties of the system. Usually, this state results when the forward reaction proceeds at the same rate as the reverse reaction.

To understand it better let's move to the principles of the state. ·The chemical equilibrium state indicates different inner reactions within the closed system that is not changed by time. If to put the words into the shell that is the correlation of the rate of the forward reaction and the rate of the reverse reaction. The \rightleftharpoons symbol denotes that phenomenon. We can get the equilibrium state only in the closed system.

We mean that such experiment will be successful just in such conditions. There are moments when some materials can lose their property because of the wrong environment. When the concentrations of reactants and products have become constant, an equation is said to have reached a point of equilibrium. Such aspects as color, pressure and other physical properties are the symbols of the equilibrium. That state is used to be named as the dynamic because the reaction is constantly in motion.

That connection does not mean that the response is over; it means that two results are balanced. The products quantity is consistent and cannot be changed. In chemistry, you can find so many exciting things, which only

staying with us you'll see yourself being in love with it. Hope you find here what you searched for.