

Neuromarketing

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Also, the slope of Freundlich Isotherm $\ln \log$ is decreasing, the graph of frequency Isotherm $\ln g$ and \log are increasing. It means that I made some mistake on measuring the I-gram dosages of activated carbon. Abstract Introduction Page 3 Experimental Apparatus and Procedure Result and Discussion and Conclusions Page 1 1 Page 1 Page 3 Theory Page 4 Page 5 Summary Page 1 1 References The objective of this experiment was to understand the properties of granular activated carbon (GAC) as an adsorbent and develop a Freundlich Isotherm.

Adsorption can involve the enterprise accumulation or concentration of substances on the surface of solids, whereas absorption is the penetration of the enterprise accumulation or concentration of substances into the solids. There are three types of adsorption, Physical Adsorption, Chemical Adsorption and Exchange Adsorption. Theory Physical adsorption is primarily due to van der Waals forces. An example of physical adsorption is the adsorption by activated carbon. Chemical adsorption represents the adsorbent undergoes chemical interaction with the adsorbent, the phenomenon is referred to as chemical adsorption or chemisorptions.

Exchange adsorption is a process in which ions of one substance concentrate at a surface. The following equations were used in this experiment. Be
Experimental Apparatus and Procedure Experimental apparatus 1) Stirrer Set 2) Beaker 3) pH meter 4) pipette 5) Cylinder 6) Spectrophotometer 7) Reagent and sample (Methyl Blue, Activated Carbon) Experimental Procedure Before starting the experiment, sample with 100 mg/L concentration of Methyl Blue solution was prepared. The initial adsorption was read with

spectrophotometer. Mill well-mixed sample was transferred into three beakers.

Then, gram usages of activated carbon were added. After adding activated carbon, the stirrer set was turned on about irresolution per minutes. The absorption of each samples were read at 10 minute, 30 minute, minute and eminent for each trial. A sample from each beaker was pulled out and put into plastic cylinder. These cylinders were placed into spectrophotometer and determined the adsorption for each beaker. These values were recorded at the end. Result Table 8. 1: initial date of the sample Concentration, MGM/L 20 Absorbent 2. 09 Graph 8. 1 the graph of Concentration Vs.. Absorbent Table 8.

The slope is supposed to be increasing. The g and log are acceptable because the slopes are increasing. There are some possible errors on measuring the data. The error possibly came from human error when we were mixing the sample. The volume of mixture maybe was not mill. Another error maybe came from the measurement. Maybe we read the value incorrectly. The improvement of this experiment is to use other techniques of measuring the volume of well-mixed sample. It can make sure the volume more precise and accurate. Summary and Conclusion Based on the data and observations, this experiment is successful.

The slope of Friendlier Isotherm in g and log are increasing. It means the values are acceptable. The slope of Friendlier Isotherm in leg is decreasing. It means the value is not acceptable. Although this experiment has an error and misreading, the experiment was done correctly and the measuring

values were acceptable. Throughout this experiment, the Freundlich Isotherm is determined. It is important for environmental engineers to understand Granulated activated carbon. References DRP. Mahout Aimed (2014). Environmental engineering laboratory manual, New York University.