Elements and their types with electron dot structures

Contents

H He

Hydrogen

- Atomic Number: 1
- Group: 1
- Period: 1
- Series: Nonmetals
- Number of Electrons (with no charge): 1
- Number of Neutrons (most common/stable nuclide): 0
- Number of Protons: 1

[image]

[image]

Li

Helium

- Atomic Number: 2
- Group: 18
- Period: 1
- Series: Noble Gasses
- Number of Electrons (with no charge): 2
- Number of Neutrons (most common/stable nuclide): 2
- Number of Protons: 2

[image]

[image]

Lithium

- Atomic Number: 3
- Group: 1
- Period: 2
- Series: Alkali Metals
- Number of Electrons (with no charge): 3
- Number of Neutrons (most common/stable nuclide): 4
- Number of Protons: 3

Be

[image]

[image]

Beryllium

- Atomic Number: 4
- Group: 2
- Period: 2
- Series: Alkali Earth Metals
- Number of Electrons (with no charge): 4
- Number of Neutrons (most common/stable nuclide): 5
- Number of Protons: 4

B

[image]

[image]

C

Boron

- Atomic Number: 5
- Group: 13
- Period: 2
- Series: Metalloids (Nonmetal)
- Number of Electrons (with no charge): 5
- Number of Neutrons (most common/stable nuclide): 6
- Number of Protons: 5

[image]

[image]

Carbon

N

- Atomic Number: 6

- Group: 14
- Period: 2
- Series: Nonmetals
- Number of Electrons (with no charge): 6
- Number of Neutrons (most common/stable nuclide): 6
- Number of Protons: 6

[image]

[image]

O

Nitrogen

- Atomic Number: 7
- Group: 15
- Period: 2
- Series:

Nonmetals

- Number of Electrons (with no charge): 7
- Number of Neutrons (most common/stable nuclide): 7
- Number of Protons: 7

[image]

[image]

F

Oxygen

- Atomic Number: 8
- Group: 16
- Period: 2
- Series: Nonmetals
- Number of Electrons

(with no charge): 8

- Number of Neutrons (most common/stable nuclide): 8
- Number of Protons: 8

[image]

[image]

Ne

Fluorine

- Atomic Number: 9
- Group: 17
- Period: 2
- Series: Halogens
- Number of Electrons (with no charge): 9
- Number of

Neutrons (most common/stable nuclide): 10

- Number of Protons: 9

[image]

[image]

Na

Neon

- Atomic Number: 10
- Group: 18
- Period: 2
- Series: Noble Gasse
- Number of Electrons (with no charge): 10
- Number of Neutrons (most common/sta

ble nuclide):

10

- Number of

Protons: 10

[image]

[image]

Mg

Sodium

Overview of Sodium

- Atomic

Number: 11

- Group: 1
- Period: 3
- Series:

Alkali

Metals

- Number of

Electrons

(with no

charge): 11

- Number of

Neutrons

(most

common/sta

12

- Number of Protons: 12

[image]

[image]

Si

Aluminum

- Atomic Number: 13
- Group: 13
- Period: 3
- Series: Metals
- Number of Electrons (with no charge): 13
- Number of Neutrons (most common/stable nuclide): 14
- Number of

Protons: 13

[image]

[image]

P

Silicon

- Atomic Number: 14
- Group: 14
- Period: 3
- Series: Metalloids (Nonmetal)
- Number of Electrons (with no charge): 14
- Number of Neutrons (most common/stable nuclide): 14
- Number of Protons: 14

[image]

[image]

S

Phosphorus

- Atomic Number: 15
- Group: 15
- Period: 3
- Series: Nonmetals
- Number of Electrons (with no charge): 15
- Number of Neutrons (most common/stable nuclide): 16
- Number of Protons: 15

[image]

[image]

Sulfur

- Atomic Number: 16
- Group: 16
- Period: 3
- Series: Nonmetals
- Number of Electrons (with no charge): 16 Cl
- Number of Neutrons (most common/stable nuclide): 16
- Number of Protons: 16

[image]

[image]

Number: 18

- Group: 18
- Period: 3
- Series:

  Noble

  Gasses
- Number of

  Electrons

  (with no

  charge): 18
- Number of

  Neutrons

  (most

  common/sta

  ble nuclide):

  22
- Number of

  Protons: 18

[image]

[image]

Ca

Potassium

- Atomic

  Number: 19

- Group: 1
- Period: 4
- Series:

  Alkali

  Metals
- Number of

  Electrons

  (with no

  charge): 19
- Number of

  Neutrons

  (most

  common/sta

  ble nuclide):

  20
- Number of

  Protons: 19

[image]

[image]

Sc

Calcium

- Atomic

  Number: 20
- Group: 2

- Period: 4
- Series:

  Alkali Earth

  Metals
- Number of

  Electrons

  (with no

  charge): 20
- Number of

  Neutrons

  (most

  common/sta

  ble nuclide):

  20
- Number of

  Protons: 20

[image]

[image]

Ti

Scandium

- Atomic

  Number: 21
- Group: 3
- Period: 4
- Series:

Transition Metals

- Number of Electrons (with no charge): 21
- Number of Neutrons (most common/stable nuclide): 24
- Number of Protons: 21

[image]

[image]

V

Titanium

- Atomic Number: 22
- Group: 4
- Period: 4
- Series: Transition

Electrons (with no charge): 23

- Number of Neutrons (most common/stable nuclide): 28
- Number of Protons: 23

[image]

[image]

Mn

Chromium

- Atomic Number: 24
- Group: 6
- Period: 4
- Series: Transition Metals
- Number of Electrons (with no

charge): 24

- Number of Neutrons (most common/stable nuclide): 28
- Number of Protons: 24

[image]

[image]

Fe

Manganese

- Atomic Number: 25
- Group: 7
- Period: 4
- Series: Transition Metals
- Number of Electrons (with no charge): 25
- Number of

Neutrons (most common/stable nuclide): 30

- Number of Protons: 25

[image]

[image]

Co

Iron

- Atomic Number: 26
- Group: 8
- Period: 4
- Series: Transition Metals
- Number of Electrons (with no charge): 26
- Number of Neutrons (most

common/sta

ble nuclide):

30

- Number of

Protons: 26

[image]

[image]

Ni

Cobalt

- Atomic

Number: 27

- Group: 9
- Period: 4
- Series:

Transition

Metals

- Number of

Electrons

(with no

charge): 27

- Number of

Neutrons

(most

common/sta

ble nuclide):

Protons: 28

[image]

Copper

- Atomic Number: 29
- Group: 11
- Period: 4
- Series: Transition Metals
- Number of Electrons (with no charge): 29
- Number of Neutrons (most common/stable nuclide): 35
- Number of Protons: 29

Zn

[image]

Zinc

- Atomic Number: 30
- Group: 12
- Period: 4
- Series: Transition Metals
- Number of Electrons (with no charge): 30
- Number of Neutrons (most common/stable nuclide): 35
- Number of Protons: 30

Ga

[image]

[image]

Ge

Gallium

- Atomic Number: 31
- Group: 13
- Period: 4
- Series: Metals
- Number of Electrons (with no charge): 31
- Number of Neutrons (most common/stable nuclide): 39
- Number of Protons: 31

[image]

[image]

Germanium

As

- Atomic Number: 32

- Group: 14
- Period: 4
- Series: Metalloids (Metals)
- Number of Electrons (with no charge): 32
- Number of Neutrons (most common/stable nuclide): 41
- Number of Protons: 32

[image]

[image]

Se

Arsenic

- Atomic Number: 33
- Group: 15

Nonmetals

- Number of Electrons (with no charge): 34
- Number of Neutrons (most common/stable nuclide): 45
- Number of Protons: 34

[image]

[image]

Kr

Bromine

- Atomic Number: 35
- Group: 17
- Period: 4
- Series: Halogens
- Number of Electrons

(with no charge): 35

- Number of Neutrons (most common/stable nuclide): 45
- Number of Protons: 35

[image]

[image]

Rb

Krypton

- Atomic Number: 36
- Group: 18
- Period: 4
- Series: Noble Gasses
- Number of Electrons (with no

charge): 36

- Number of Neutrons (most common/stable nuclide): 48
- Number of Protons: 36

[image]

[image]

Sr

Rubidium

- Atomic Number: 37
- Group: 1
- Period: 5
- Series: Alkali Metals
- Number of Electrons (with no charge): 37
- Number of

Neutrons (most common/stable nuclide): 48

- Number of Protons: 37

[image]

[image]

Y

Strontium

- Atomic Number: 38
- Group: 2
- Period: 5
- Series: Alkali Earth Metals
- Number of Electrons (with no charge): 38
- Number of Neutrons

common/sta

ble nuclide):

50

- Number of

Protons: 39

[image]

[image]

Nb

Zirconium

- Atomic

Number: 40

- Group: 4
- Period: 5
- Series:

Transition

Metals

- Number of

Electrons

(with no

charge): 40

- Number of

Neutrons

(most

common/sta

ble nuclide):

51

- Number of Protons: 40

[image]

[image]

Mo

Niobium

- Atomic Number: 41
- Group: 5
- Period: 5
- Series: Transition Metals
- Number of Electrons (with no charge): 41
- Number of Neutrons (most common/stable nuclide): 52
- Number of

Protons: 41

[image]

[image]

Molybdenum Tc

- Atomic Number: 42
- Group: 6
- Period: 5
- Series: Transition Metals
- Number of Electrons (with no charge): 42
- Number of Neutrons (most common/stable nuclide): 54
- Number of Protons: 42

[image]

[image]

Ru

Technetium

- Atomic Number: 43
- Group: 7
- Period: 5
- Series: Transition Metals
- Number of Electrons (with no charge): 43
- Number of Neutrons (most common/stable nuclide): 55
- Number of Protons: 43

[image]

Rhodium

- Atomic Number: 45
- Group: 9
- Period: 5
- Series: Transition Metals
- Number of Electrons (with no charge): 45
- Number of Neutrons (most common/stable nuclide): 58

Pd

Number of Protons: 45

[image]

Ag

Palladium

- Atomic

Number: 46

- Group: 10
- Period: 5
- Series: Transition Metals
- Number of Electrons (with no charge): 46
- Number of Neutrons (most common/stable nuclide): 60
- Number of Protons: 46

[image]

Cd

Silver

- Atomic Number: 47
- Group: 11

- Period: 5
- Series: Transition Metals
- Number of Electrons (with no charge): 47
- Number of Neutrons (most common/stable nuclide): 61
- Number of Protons: 47

[image]

Cadmium

In

- Atomic Number: 48
- Group: 12
- Period: 5
- Series: Transition

Metals

- Number of Electrons (with no charge): 48
- Number of Neutrons (most common/stable nuclide): 64
- Number of Protons: 48

[image]

[image]

Sn

Indium

- Atomic Number: 49
- Group: 13
- Period: 5
- Series: Metals
- Number of Electrons

Neutrons (most common/stable nuclide): 69

- Number of Protons: 50

[image]

[image]

Te

Antimony

- Atomic Number: 51
- Group: 15
- Period: 5
- Series: Metalloids (Metals)
- Number of Electrons (with no charge): 51
- Number of Neutrons (most

common/sta

ble nuclide):

71

- Number of

Protons: 51

[image]

[image]

I

Tellurium

- Atomic

Number: 52

- Group: 16
- Period: 5
- Series:

Metalloids

(Nonmetal)

- Number of

Electrons

(with no

charge): 52

- Number of

Neutrons

(most

common/sta

ble nuclide):

76

- Number of Protons: 52

[image]

[image]

Xe

Iodine

- Atomic Number: 53
- Group: 17
- Period: 5
- Series: Halogens
- Number of Electrons (with no charge): 53
- Number of Neutrons (most common/stable nuclide): 74
- Number of

Protons: 53

[image]

[image]

Cs

Xenon

- Atomic Number: 54
- Group: 18
- Period: 5
- Series: Noble Gasses
- Number of Electrons (with no charge): 54
- Number of Neutrons (most common/stable nuclide): 77
- Number of Protons: 54

[image]

Barium

- Atomic Number: 56
- Group: 2
- Period: 6
- Series: Alkali Earth Metals
- Number of Electrons (with no charge): 56

La

- Number of Neutrons (most common/stable nuclide): 81
- Number of Protons: 56

[image]

[image]

Lanthanum

- Atomic Number: 57
- Group: 3
- Period: 6
- Series: Lanthanides
- Number of Electrons (with no charge): 57
- Number of Neutrons (most common/stable nuclide): 82
- Number of Protons: 57

Hf

[image]

[image]

Ta

Hafnium

- Atomic

Number: 72

- Group: 4
- Period: 6
- Series: Transition Metals
- Number of Electrons (with no charge): 72
- Number of Neutrons (most common/stable nuclide): 106
- Number of Protons: 72

[image]

[image]

Ta

Tantalum

- Atomic Number: 73

- Group: 5
- Period: 6
- Series: Transition Metals
- Number of Electrons (with no charge): 73
- Number of Neutrons (most common/stable nuclide): 108
- Number of Protons: 73

[image]

[image]

W

Tantalum

- Atomic Number: 73
- Group: 5

Transition Metals

- Number of Electrons (with no charge): 74
- Number of Neutrons (most common/stable nuclide): 110
- Number of Protons: 74

[image]

[image]

Ir

Osmium

- Atomic Number: 76
- Group: 8
- Period: 6
- Series: Transition

Metals

- Number of Electrons (with no charge): 76
- Number of Neutrons (most common/stable nuclide): 114
- Number of Protons: 76

[image]

[image]

Pt

Iridium

- Atomic Number: 77
- Group: 9
- Period: 6
- Series: Transition Metals
- Number of

Electrons (with no charge): 77

- Number of Neutrons (most common/stable nuclide): 115
- Number of Protons: 77

[image]

Au

Platinum

- Atomic Number: 78
- Group: 10
- Period: 6
- Series: Transition Metals
- Number of Electrons (with no

charge): 78

- Number of Neutrons (most common/stable nuclide): 117
- Number of Protons: 78

[image]

Hg

Gold

- Atomic Number: 79
- Group: 11
- Period: 6
- Series: Transition Metals
- Number of Electrons (with no charge): 79
- Number of Neutrons

121

- Number of Protons: 80

[image]

[image]

Pb

Thallium

- Atomic Number: 81
- Group: 13
- Period: 6
- Series: Metals
- Number of Electrons (with no charge): 81
- Number of Neutrons (most common/stable nuclide): 123
- Number of

Protons: 81

[image]

[image]

Bi

Lead

- Atomic Number: 82
- Group: 14
- Period: 6
- Series: Metals
- Number of Electrons (with no charge): 82
- Number of Neutrons (most common/stable nuclide): 125
- Number of Protons: 82

[image]

[image]

Bismuth

- Atomic Number: 83
- Group: 15
- Period: 6
- Series: Metals
- Number of Electrons (with no charge): 83 Po
- Number of Neutrons (most common/stable nuclide): 126
- Number of Protons: 83

[image]

[image]

Polonium

- Atomic Number: 84
- Group: 16
- Period: 6
- Series: Metalloids (Metals)
- Number of Electrons (with no charge): 84
- Number of Neutrons (most common/stable nuclide): 125
- Number of Protons: 84

[image]

[image]

Astatine

- Group: 18
- Period: 6
- Series:

  Noble

  Gasses
- Number of

  Electrons

  (with no

  charge): 86
- Number of

  Neutrons

  (most

  common/sta

  ble nuclide):

  136
- Number of

  Protons: 86

[image]

[image]

Ra

Francium

- Atomic

  Number: 87
- Group: 1

- Period: 7
- Series:

  Alkali

  Metals

- Number of

  Electrons

  (with no

  charge): 87
- Number of

  Neutrons

  (most

  common/sta

  ble nuclide):

  136
- Number of

  Protons: 87

[image]

[image]

Ac

Radium

- Atomic

  Number: 88
- Group: 2
- Period: 7

- Series:

  Alkali Earth

  Metals
- Number of

  Electrons

  (with no

  charge): 88
- Number of

  Neutrons

  (most

  common/sta

  ble nuclide):

  138
- Number of

  Protons: 88
- [image]

[image]

Actinium

- Atomic

  Number: 89
- Group: 3
- Period: 7
- Series:

Actinides

- Number of Electrons (with no charge): 89
- Number of Neutrons (most common/stable nuclide): 138
- Number of Protons: 89

[image]

[image]