

6 strong acids and 6 strong bases flashcard



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- Examples of the superbases include:

Strong bases are bases which completely dissociate in solution to yield hydroxide ions, or deprotonate water to yield hydroxide ions OH^- (hydroxide ion). The hydroxides of the Group I (alkali metals) and Group II (alkaline earth) metals usually are considered to be strong bases. These are classic Arrhenius bases. Here is a list of the most common strong bases.

The list of 6 Strong Acids and bases:

Hydrochloric Acid (HCl)(Strong Acid)Hydrobromic Acid (HBr)(Strong Acid)Hydroiodic Acid (HI)(Strong Acid)Nitric Acid (HNO_3)(Strong Acid)Perchloric Acid (HClO_4)(Strong Acid)Sulfuric Acid (H_2SO_4)(Strong Acid)Lithium Hydroxide (LiOH)(Strong Base)Sodium Hydroxide (NaOH)(Strong Base)Potassium Hydroxide (KOH)(Strong Base)Calcium Hydroxide ($\text{Ca}(\text{OH})_2$)(Strong Base)Strontium Hydroxide ($\text{Sr}(\text{OH})_2$)(Strong Base)Barium Hydroxide ($\text{Ba}(\text{OH})_2$)(Strong Base)

Examples of the superbases include:

- Ethoxide ion
- Butyl lithium ($n\text{-BuLi}$)
- Lithium diisopropylamide (LDA) ($\text{C}_6\text{H}_{14}\text{LiN}$)
- Lithium diethylamide (LDEA)
- Sodium amide (NaNH_2)
- Sodium hydride (NaH)
- Lithium bis(trimethylsilyl)amide, $((\text{CH}_3)_3\text{Si})_2\text{NLi}$