

Chemical reactions

[Science](#), [Chemistry](#)



Chemical Reactions: Combustion: Lab Worksheet and Rubric Before You Begin: You may either copy and paste this document into a word processing program of your choice or print this page. Procedure:

1. Iron (III) and copper (II) sulfate solution Fill a small test tube halfway with copper (II) sulfate solution. Add a 2.0 gram iron rod to the solution and observe the reaction.

2. Lead (II) nitrate and potassium iodide solutions Pour about 2.0 mL of lead (II) nitrate into the test tube. Add 5 to 10 drops of potassium iodide solution to the test tube and record your observations of the reaction.

3. Magnesium metal and hydrochloric acid solution Place one scoop of magnesium turnings into the test tube. Add hydrochloric acid slowly and observe the reaction. Next, place a burning splint near the mouth of the test tube to test for the presence of hydrogen gas.

4. Electrolysis of water Use a U-tube with electrodes at each end, connected to a battery. Fill the U-tube with water. Turning on the battery, observe the results at each electrode. Next, test for hydrogen and oxygen gas produced at each end of the tube.

5. Burning magnesium Place magnesium ribbon in a clean crucible (on a clay triangle above a Bunsen burner). Heat until the magnesium begins to burn.

Data and Observations: Complete a data table that includes a prediction of reaction type (single replacement, double replacement, synthesis, decomposition, or combustion), observations, and identification of reaction type for each reaction in the lab.