

# Chemical physical properties essay



**ASSIGN  
BUSTER**

Min- almost all the way melted maybe a quarter of the cube left. 50 min- just about all gone small pieces remain.

57 min ice is completely melted. Table 2: Freezing Point Data 6: 07 pm 7: 37 pm 30 min- No change 60 min- partially frozen has ice crystals 90 min- frozen 6: 07 pr. 8: 07 pm 60 min- has very little ice forming 90 min- partially frozen still watery 120 min- mostly frozen Post-lab Questions 1 . Did both ice cubes melt at the same rate? Why or why not? No the ice in tub B with the salt melted considerably faster. 2. Did both ice cubes re-freeze at the same time? Why or why not? No the one in tub A (without salt) froze well before the one with salt.

This is because the salt caused the freezing point to be lower in the water. 3. Is the freezing point of water a chemical or physical property? 4. What effect did the salt have on the ice cubes? It caused the ice cube in tub B to melt considerably faster than the cube in tub A. 5.

Is the solubility of table salt (NCAA) a chemical or physical property? 6. Why might some people put salt on ice during the winter? TO melt it so its not as slippery. 7. Is the solubility of table salt (NCAA) a chemical or physical property? Experiment 1: Disruption of Micelles Table 3: Reaction Rate Data and Observations Test Tube Solution Components Original Observations (Time 0 minutes) Final Observations (Time = 30 minutes) Water and powdered milk White and milky No change stayed looking like milk Water, powdered milk, and vinegar Separated at the top Separation and curdling moved about one third of the solution C Water powered milk, vinegar (stirred) Separating throughout the entire test tube. Looks like chunks Entire

test tube curdled Water powdered milk, soap Soap sitting on top of solution  
Water, powdered milk, soap (stirred) Yellowish milky color throughout the  
tube Looks as if soap took over the milk in the tube. 1. What was the effect  
of the acetic acid on the milk? It caused the milk to curdle. 2.

What differences, if any, did you notice between test tubes B and C? The  
biggest difference was that in tube C the milk curdles throughout the entire  
test tube well it only curdled on the top in tube B. 3. What was the effect of  
the dish soap on the milk? The soap took over the milk. 4. Was the dish soap  
more or less dense than the milk? How do you know? The soap was less  
dense than the milk because it sat on top of the milk solution the entire time.